

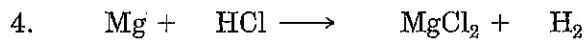
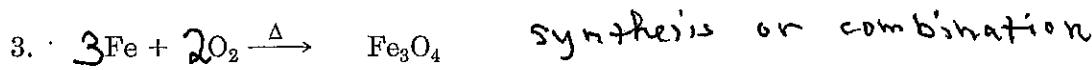
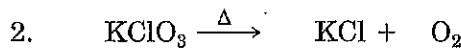
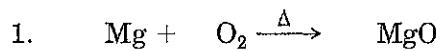
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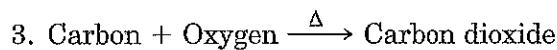
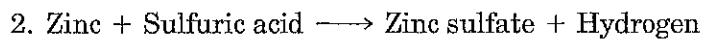
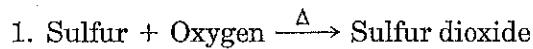
INSTRUCTOR \_\_\_\_\_

**EXERCISE 6****Equation Writing and Balancing I**

*Name the type (9.3)*  
 Balance the following equations:



Beneath each word equation write the formula equation and balance it. Remember that oxygen and hydrogen are diatomic molecules.



## EXERCISE 6 (continued)

NAME \_\_\_\_\_

Balance the following equations:

1.  $\text{N}_2 + \text{H}_2 \xrightarrow{\Delta} \text{NH}_3$
2.  $\text{CoCl}_2 \cdot 6\text{H}_2\text{O} \xrightarrow{\Delta} \text{CoCl}_2 + 6\text{H}_2\text{O}$
3.  $\text{Fe} + \text{H}_2\text{O} \xrightarrow{\Delta} \text{Fe}_3\text{O}_4 + \text{H}_2$
4.  $\text{F}_2 + \text{H}_2\text{O} \xrightarrow{\Delta} \text{HF} + \text{O}_2$
5.  $\text{Pb}(\text{NO}_3)_2 \xrightarrow{\Delta} \text{PbO} + \text{NO} + \text{O}_2$

Beneath each word equation write and balance the formula equation. Oxygen, hydrogen, and bromine are diatomic molecules

